

Chemistry Part II Test Review

1. List the five signs of a chemical change/reaction:

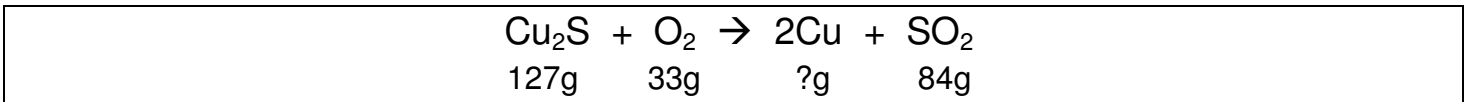
- a. Please _____
- b. Excuse _____
- c. Coughs _____
- d. Sneezes _____
- e. Burps _____

What is a precipitate? _____

2. Circle the chemical reactions.

- Cutting the grass dissolving a tablet in water (that creates bubbles) boiling water
- Baking a cake burning sugar tearing paper
- A nail rusting mixing Kool-Aid scraping rust off of a bike

- 3. What does the Law of Conservation of Mass (Matter) say? _____
- 4. _____



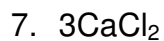
Using the reaction above – According to the Law of Conservation of Mass, what is the mass of Cu?
 _____g

- 5. What information is found in a chemical formula? _____
- _____

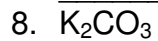
Count the atoms in each of the following formulas:

6. Na₂CrO₄

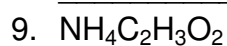
Element	# of Atoms	Total molecules	Total atoms
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____



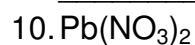
Element	# of Atoms	Total molecules	Total atoms
_____	_____	_____	_____
_____	_____		
_____	_____		
_____	_____		



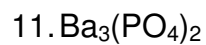
Element	# of Atoms	Total molecules	Total atoms
_____	_____	_____	_____
_____	_____		
_____	_____		
_____	_____		



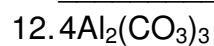
Element	# of Atoms	Total molecules	Total atoms
_____	_____	_____	_____
_____	_____		
_____	_____		
_____	_____		



Element	# of Atoms	Total molecules	Total atoms
_____	_____	_____	_____
_____	_____		
_____	_____		
_____	_____		



Element	# of Atoms	Total molecules	Total atoms
_____	_____	_____	_____
_____	_____		
_____	_____		
_____	_____		



Element	# of Atoms	Total molecules	Total atoms
_____	_____	_____	_____
_____	_____		
_____	_____		
_____	_____		

13. What is the arrow sign called in a chemical equation? _____

14. What does it mean? _____

15. What are the substances on the left of the arrow in a chemical equation called?

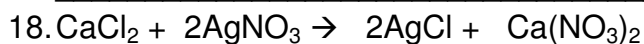
16. What are the substances on the right of the arrow in a chemical equation called?

All: answer the following questions about each reaction

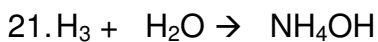
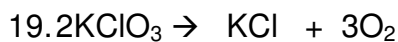
OL: Circle the following equations that follow the Law of Conservation of Matter/Mass. (balanced)

HONORS: If a reaction is not balanced, write the correct balanced equation. If it is not possible, write "not possible".

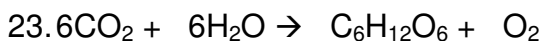
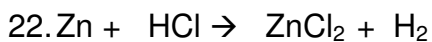
17. What is a balanced equation? What does balanced mean in this case? _____



a. How many reactants are there in the equation above?

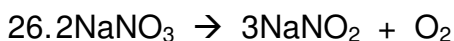
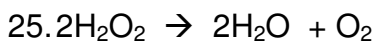


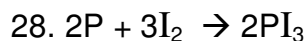
a. How many COMPOUNDS are in the equation above? _____



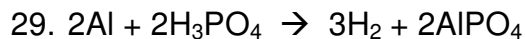
a. How many atoms of Oxygen are on the reactants side of the equation? _____

b. How many atoms of Oxygen are on the products side of the equation? _____





a. How many elements are there in the equation above? _____



a. How many molecules in the reactants of the equation above? _____

b. How many molecules in the products of the equation above? _____

Chemistry Part I Test Review

30. What are the three subatomic particles? _____

31. How do you find the number of protons in an atom? _____

32. How do you find the number of electrons in an atom? _____

33. How do you find the number of neutrons in an atom? _____

34. Complete the following table about the three subatomic particles.

Name	Charge	Location in atom	Mass of particle	Symbol in models
Proton				
	negative			
				n^0

35. Which subatomic particles are the heaviest? _____

36. Which one is the lightest? _____

37. Which subatomic particle determines the identity of an atom? _____

38. Which subatomic particles contribute mass to an element? _____

39. The majority of the mass of an atom is found in its _____

40. Does altering the number of electrons change the element? _____

41. Does altering the number of protons change the element? _____

Define the following terms:

42. Element: _____

43. Compound: _____

44. Mixture: _____

45. Rows on the Periodic Table are called: _____ and run _____

46. Columns on the Periodic Table are called _____ and run _____

47. Where are the metals located on the Periodic Table? _____

48. Where are nonmetals located on the periodic table? _____

49. Where are metalloids located on the Periodic Table? _____

50. Define metalloids: _____

51. Define energy levels: _____

52. How can the number of energy levels an atom has be determined by looking at the Periodic Table? _____

53. Give an example of a chemical property _____

54. Define valence electrons: _____

55. How can the number of valence electrons be determined by looking at the Periodic Table?

56. What do elements in the same group have in common? _____

Draw Bohr models for the following elements:

57. Neon #10	58. Fluorine #9
59. Carbon #6	60. Beryllium #4