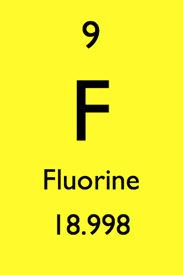
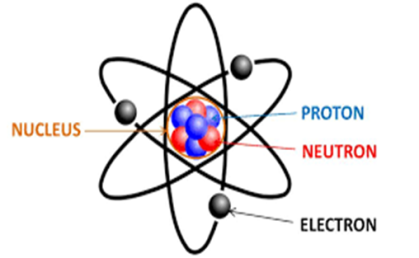
**Chemical Formulas**

* \_\_\_\_\_\_\_\_\_\_\_ are the basic units of \_\_\_\_\_\_\_\_. Atoms join together \_\_\_\_\_\_\_\_\_\_\_\_\_ to form molecules.
* Elements are pure \_\_\_\_\_\_\_\_\_\_\_\_.
* Elements are represented by 1-2 letters called \_\_\_\_\_\_\_\_\_ on the \_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_.

**Elements**

* Elements are \_\_\_\_\_\_\_ substances made from atoms of that element.

- Examples: \_\_\_ and Mg

**Molecules**

* Molecule is a term used to refer to \_\_\_\_ or \_\_\_\_\_\_ atoms which are \_\_\_\_\_\_\_\_ bonded together.

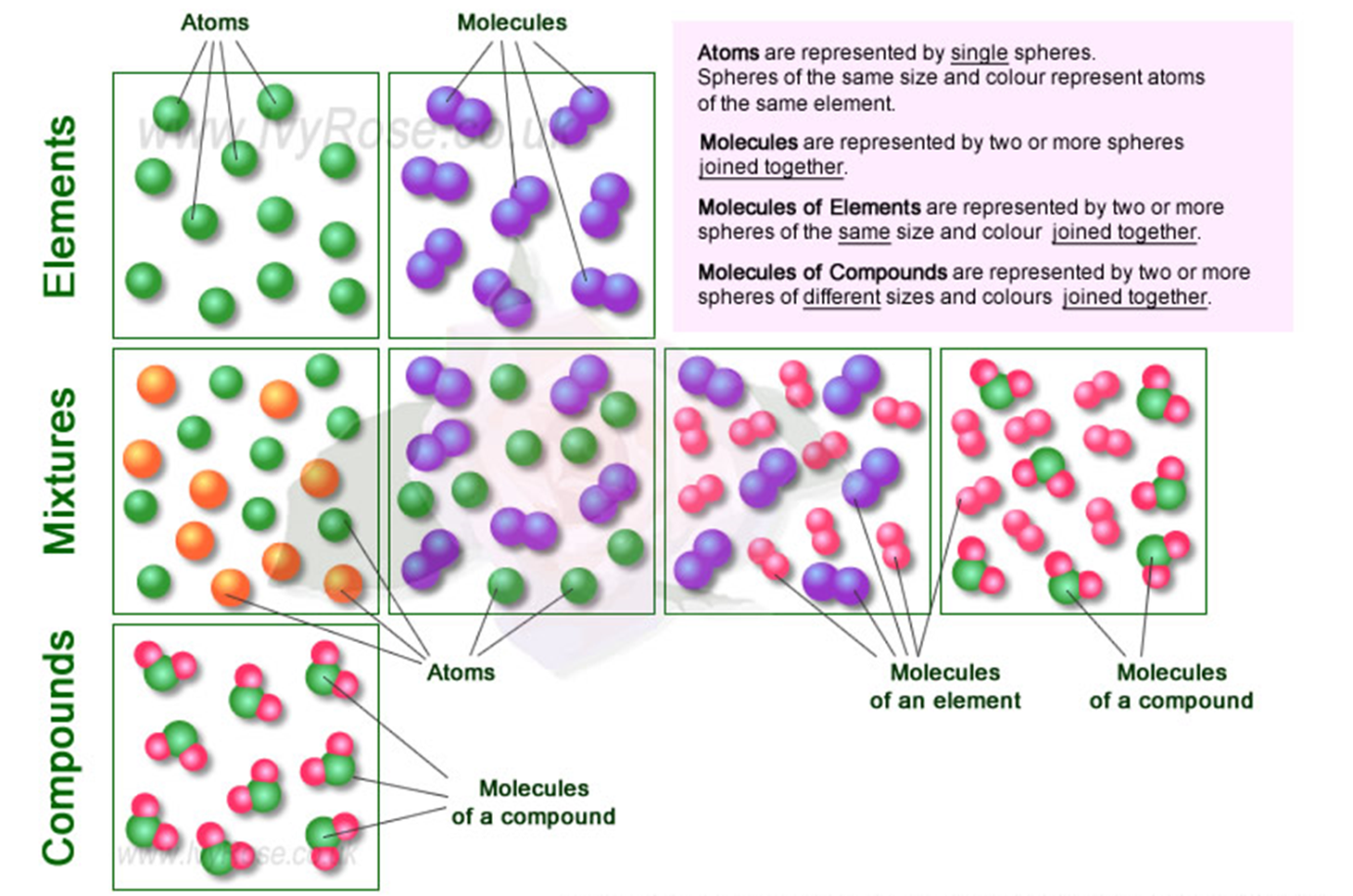
-Examples: O2 and CO2

**Compounds**

* Compounds are substances made of \_\_\_\_\_\_ or more different elements \_\_\_\_\_\_\_\_\_\_\_\_ combined in a set ratio.
* A chemical \_\_\_\_\_\_\_\_\_\_\_, or electric force holds the atoms together in a compound.
* CO2 is both a \_\_\_\_\_\_\_\_\_\_and a \_\_\_\_\_\_\_\_\_\_ as well.
* The properties of a compound are usually \_\_\_\_\_\_\_\_\_\_ from the elements that make up the compound.

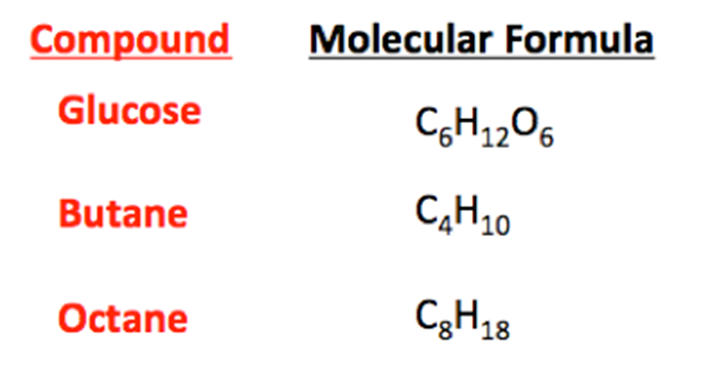
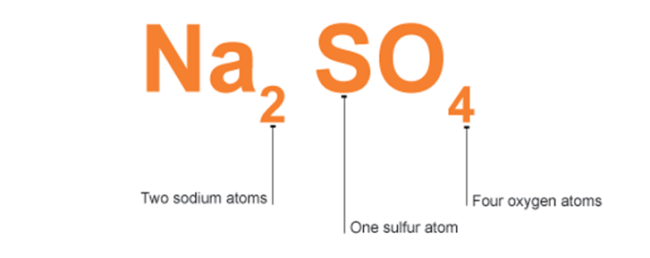
-Example

* + Salt is composed of \_\_\_\_\_\_\_\_\_\_\_ (Na); a silvery metal; and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ (Cl); a stinky green gas
    - When they bond to form \_\_\_\_\_\_\_\_\_, they form solid white crystals which you can eat!



**Chemical Formulas**

* Symbolic representation of a chemical compound.
* \_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_ are used to identify substances and determine the \_\_\_\_\_\_\_\_\_\_ of atoms of each element.
* The elements present in a compound are identified with the element’s \_\_\_\_\_\_\_\_\_\_\_\_.
* If more than \_\_\_\_\_ atom of an element is in the compound, it is shown with a \_\_\_\_\_\_\_\_\_\_\_.



**Subscripts**

* + Small numbers behind the element or compound
  + Shows the number of atoms in each substance.
* NH3 Total Atoms:

Different Elements:

* NaCl Total Atoms:

Different Elements:

* H2O2 Total Atoms:

Different Elements:

**Coefficients**

* + Numbers in front of the chemical formula
  + Shows the number of molecules of each Chemical Formula.

Total Atoms:

Different Elements:

Molecules:

Different Elements:

* 2NH3

Total Atoms:

Different Elements:

Molecules:

Different Elements:

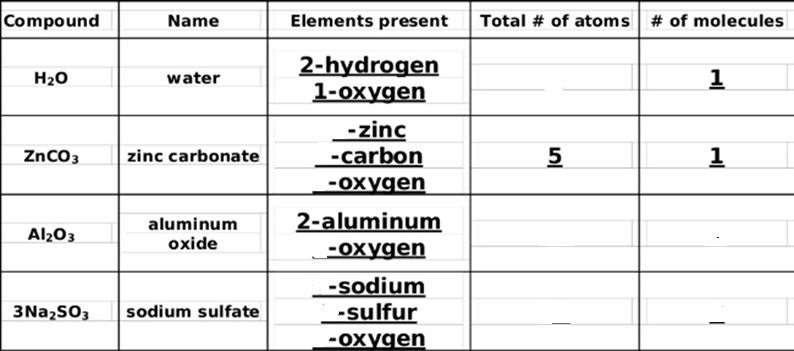
* 3NaCl
* 4H2O2

Total Atoms:

Different Elements:

Molecules:

Different Elements:

**Write the number of atoms, elements & molecules for each formula:**

1. NaOH

Total Atoms:\_\_\_\_\_\_\_  
Different Elements:\_\_\_\_\_\_\_  
Molecules:\_\_\_\_\_\_\_\_

1. MgCl2

Total Atoms:\_\_\_\_\_\_\_  
Different Elements:\_\_\_\_\_\_\_  
Molecules:\_\_\_\_\_\_\_\_

1. 2NaOH

Total Atoms:\_\_\_\_\_\_\_  
Different Elements:\_\_\_\_\_\_\_  
Molecules:\_\_\_\_\_\_\_\_

1. NaC2H3O2
2. 3HNO

Total Atoms:\_\_\_\_\_\_\_  
Different Elements:\_\_\_\_\_\_\_  
Molecules:\_\_\_\_\_\_\_\_

1. Li2SO4

Total Atoms:\_\_\_\_\_\_\_  
Different Elements:\_\_\_\_\_\_\_  
Molecules:\_\_\_\_\_\_\_\_

1. 3Al2O3

Total Atoms:\_\_\_\_\_\_\_  
Different Elements:\_\_\_\_\_\_\_  
Molecules:\_\_\_\_\_\_\_\_

1. 4Li2O

Total Atoms:\_\_\_\_\_\_\_  
Different Elements:\_\_\_\_\_\_\_  
Molecules:\_\_\_\_\_\_\_\_

Total Atoms:\_\_\_\_\_\_\_  
Different Elements:\_\_\_\_\_\_\_  
Molecules:\_\_\_\_\_\_\_\_